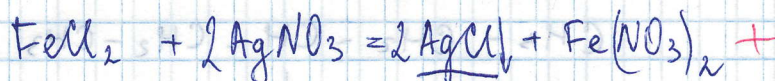


X-10-2

(98)

Указанный стан алюминия по химии.



n1 Dano:

Na

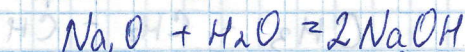
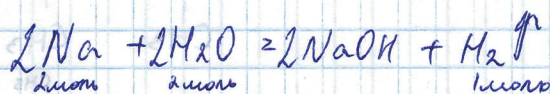
Na₂O+ H₂O

$$V(H_2) = 4,48 \text{ л}$$

$$m(p-pa) = 240 \text{ г}$$

$$w(NaOH) = 10\% = 0,1$$

$$w(Na) = ?$$



$$1. \quad n(Na) = \frac{2}{1} n(H_2) = \#$$

$$n(H_2) = \frac{V}{V_m} = \frac{4,48 \text{ л}}{22,4 \text{ л/моль}} = 0,2 \text{ моль}$$

$$n(Na) = 2 \cdot 0,2 = 0,4 \text{ моль}$$

$$m(Na) = n \cdot M = 0,4 \text{ моль} \cdot 23 = 9,2 \text{ г}$$

$$m(NaOH) = m(p-pa) \cdot w(NaOH) =$$

$$= 240 \text{ г} \cdot 0,1 = 24 \text{ г}$$

$$n(NaOH) = \frac{m}{M} = \frac{24}{40} = 0,6 \text{ моль}$$

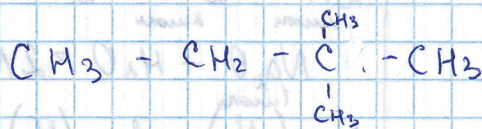
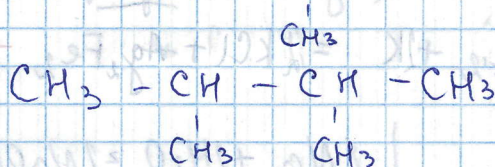
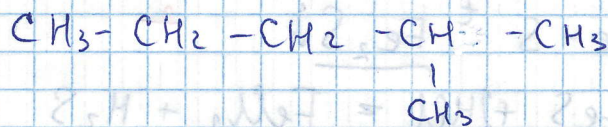
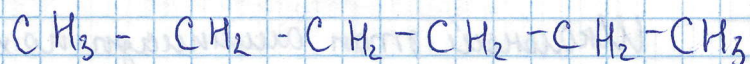
$$n(Na_2O) = \frac{1}{2} \cdot n(NaOH) = \frac{0,6}{2} = 0,3 \text{ моль}$$

$$m(Na_2O) = n \cdot M = 0,3 \cdot 62 = 18,6 \text{ г}$$

$$w(Na) = \frac{m(Na)}{m(p-pa)} = \frac{9,2}{240} = 100\% = 7,75\%$$

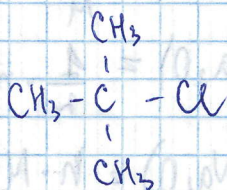
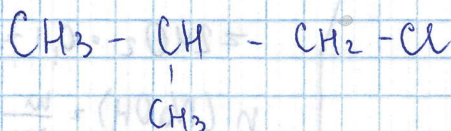
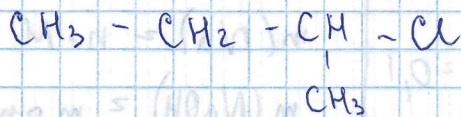
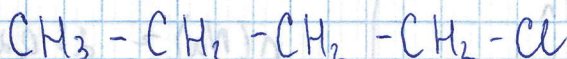
$$\text{Отв. } w(Na) = 7,75\%$$

№5. C_6H_{14} - алкан



28

C_4H_9Cl



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